

Chapter - 4 Heat

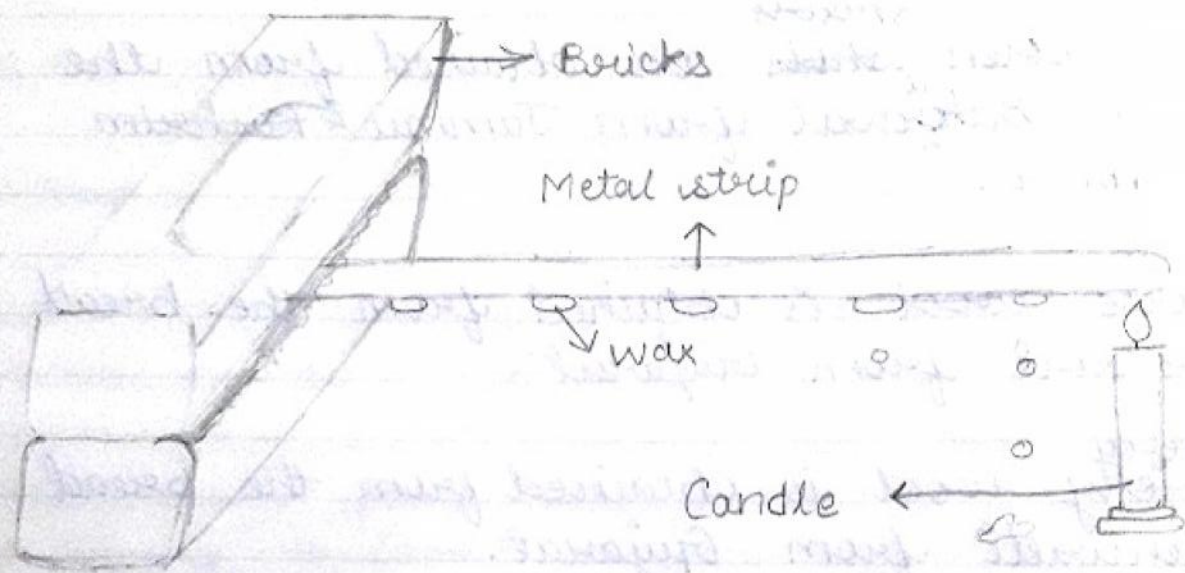
Q1. Describe a clinical thermometer. What is its function? Draw a neat and labeled diagram (fig- 4.2).

A: The thermometer that measures that measures our body temperature is called clinical thermometer. A clinical thermometer consist a long, narrow, uniform glass tube. It has a bulb at one end. This bulb contains mercury. Outside the bulb there is a small shiny thread of mercury can be seen. A clinical thermometer reads the temperature from 35°C to 42°C . It is used to measure the body temperature.

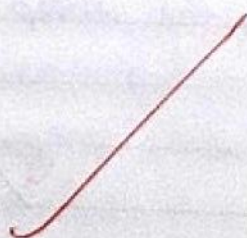
Q2. How is a clinical thermometer read?

A: To read a clinical thermometer first note the temperature ~~difference~~ indicated between the two ^{big} marks. Also note the number of divisions shown by smaller mark between these bigger marks. Suppose, the bigger ^{mark} reads 1°C and there are

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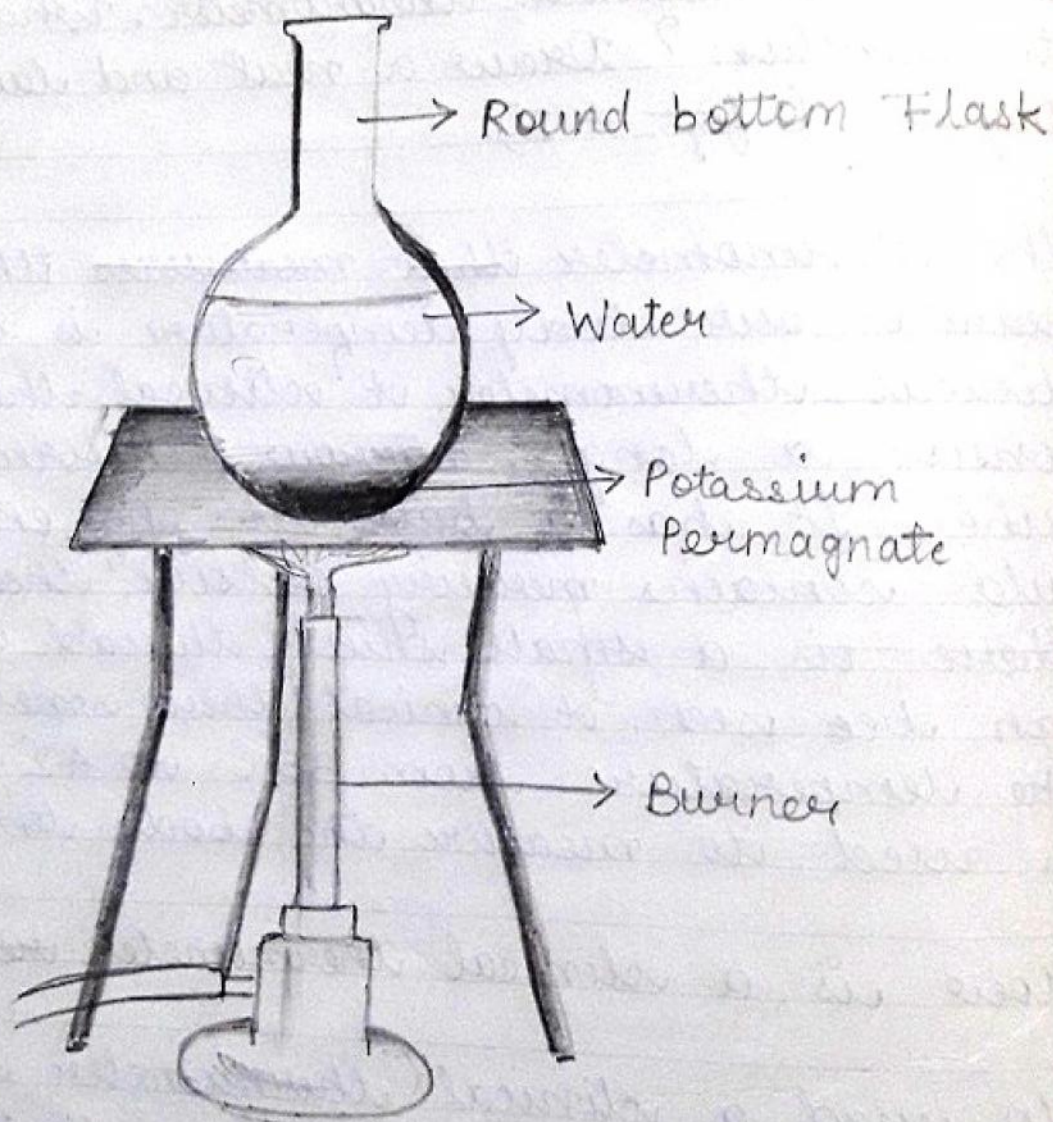
4.7 Flow of heat through a metal strip.



five divisions between them, then one small division can be read as $1/5 = 0.2^{\circ}\text{C}$

Q3. What are the precautions to be taken while handling a clinical thermometer?

- * ① Thermometer should be washed before & after use preferably with an antiseptic solution.
- ② Ensure that before use the mercury level is below 35°C .
- ③ Read the thermometer keeping the level of mercury along the line of sight.
- ④ Handle the thermometer with care. If it hits against some hard object it can break.
- ⑤ Do not hold the thermometer by the bulb while reading it.



4.9 Convection of heat in water.

Q4. State the similarity between the laboratory thermometer & a clinical thermometer.

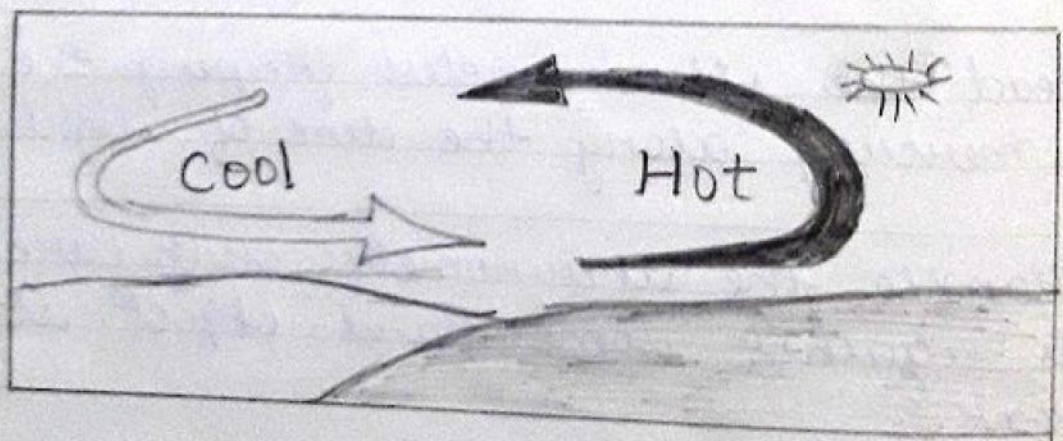
A: ① Both the thermometers have mercury in its scale.

② Both of them are used to measure the temperature.

Q5. What are the differences between the clinical thermometer & a laboratory thermometer.

Clinical Thermometer	Laboratory thermometer
① Its range is from 35°C to 42°C .	① Its range is from -10°C to 110°C .
② It is used to measure the temperature of human body.	② It is used to measure the temperature of other objects.
③ It is shorter in length than the laboratory thermometer.	③ It is longer in length than a clinical thermometer.

Day time



④ It has only mercury in its bulb.

④ It has both mercury as well as alcohol in its bulb.

⑤ It has a kink in it.

⑤ It does not have a kink.

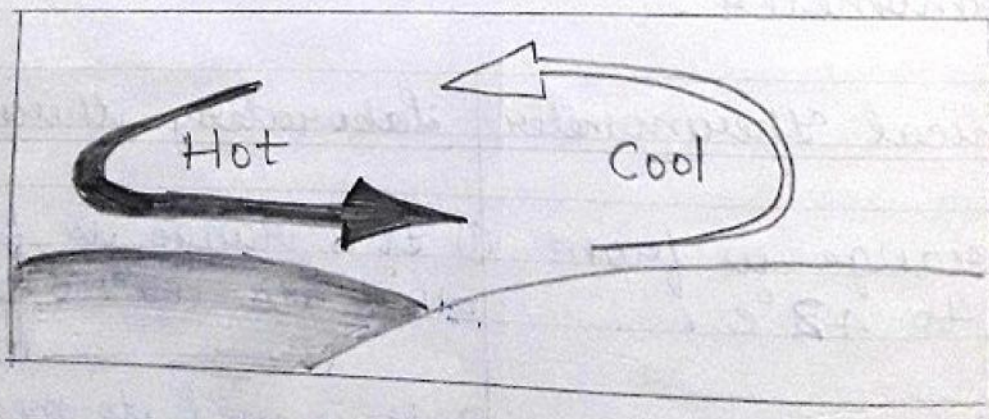
Q6. While wearing more layer of the clothes during the winter keeps us warmer than wearing just 1 thick piece of cloth? Give reason.

A: In winter we use wollen clothes. Wool is a poor conductor of heat. Moreover, air is trapped in between the wollen fibres. This air prevents the flow of heat from our body to cold surroundings, so we feel warm but if we wear more layer more air will be trapped between them and keep us more warm.

Q7. Why is it not convenient to measure the body temperature with the help of laboratory thermometer?

~~Why is it not convenient to use a~~
or
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Night time



laboratory thermometer to measure the ^{temperature} of human body?

A: It is not convenient to use laboratory thermometer to note down the temperature of the human body because, we would have to take the thermometer out of our mouth, but the laboratory thermometer should be surrounded by all the sides by the substance of which the temperature is to be measured. Otherwise, the level of mercury may fall down.

Q8. Why does not the mercury level fall or rise in a clinical thermometer when taken out?

A: A clinical thermometer has a kink which prevents the mercury from falling down. ^{Thus} ~~so~~ the mercury level does not fall in the clinical thermometer when taken out of the mouth.

Please Turn Over →

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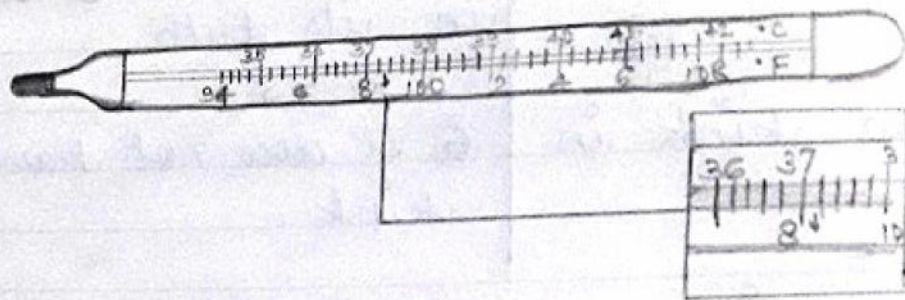


Fig 4.2 - A Clinical thermometer



4.6 - A clinical thermometer has a kink in it.



Q9. What are conductors?

A: The materials which allow heat to pass through them are called conductors.
for eg. - Iron, copper wire.

Q10. What is conduction?

A: The process by which heat is transferred from the hotter end to the cooler end of ^{an} ~~the~~ object is called conduction. for eg. - Iron, copper allows heat to pass through ~~it~~ them.

Q11. Write an activity to show conduction in solids.

A: Aim: To show conduction in solids.

* Materials Required: An iron strip, a candle and a ~~tripod~~ stand bricks.

* Method: ① Fix a few small wax pieces on the iron strip by tilting the candle at equal distance.

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Iron, copper wire, conductors of heat.

Iron, copper wire, conductors of heat.

It allows heat to pass.

- ② Fix One end of the iron strip to the stand between the two bricks.
- ③ Keep the other end of the iron strip with the candle and observe.

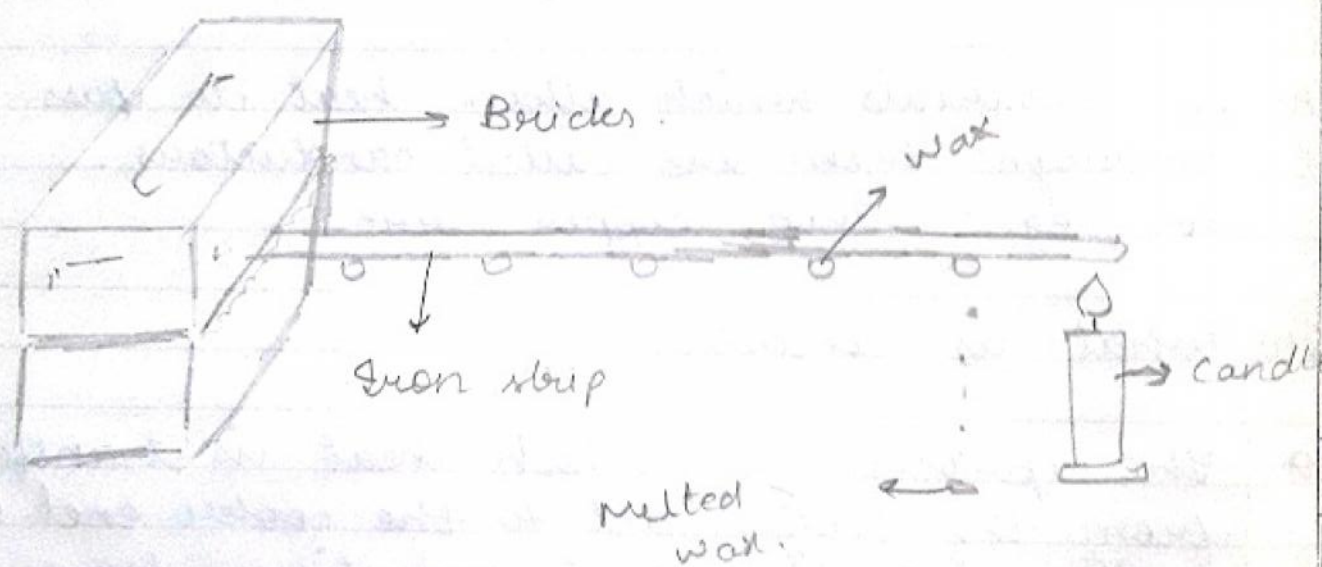
* Observation: ① Wax pieces start falling one by one. The wax pieces which is nearest to the candle fall first and the other pieces accordingly.

* Conclusion: The process by which heat is transferred from the hotter end to the colder end of an object is known as conduction. The candle first heated the nearest end and the other end afterwards. This proves the process of conduction in solids.

Q.12. Define Conduction Convection.

In liquids and gases heat is generally transferred by convection. In this process the object is made lighter in weight by heat so that it raises up. Colder objects from the sides move in & it also gets

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heated up and raises up. This process continues and whole of the object gets heated up.

Q13 Write an Activity to show convection of heat in water or liquids.

* Aim : To show convection of heat in water or liquids.

* Materials Required : A round bottom flask, a ~~tripod~~^{tripod} stand, water, burner and a crystal of potassium permanganate.

* Method : ① Fill the flask $\frac{2}{3}$ rd with water.
② Put a crystal of potassium permanganate.
③ Keep the burner just below the flask and heat it.

* Observation : When the water is heated ~~up~~ it ~~raises~~^{raises} up and moves in. This water also ~~get~~^{gets} heated & raises up. This process continues till the whole water ~~is~~^{gets} heated.

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^{to}
Conclusion : As the water gets heated and raises, the water from the sides moves in and takes ^{its place}. This proves the heating of water by convection.

Q14 In places of hot climate it is considered to paint the outer wall of the house with the white colour. Give reason.

A: In places of hot climate it is considered to paint the outer wall of the house with white colour because it reflects most of the light rays that fall on it and does not allow heat to enter into the house. Therefore, we feel cooler.

Q15. What are the different means by which heat can be transferred? Give some examples.

A: There are 3 ways by which heat can be transferred. They are as follows: -
conduction Conduction

① Conduction: The process by which heat is transferred from the hotter end to the cooler end of an object is known as

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~~conduction~~
~~convection~~ . eg :- ^{Even} ~~aluminum~~, copper allow heat to pass through them easily.

(2) Convection : When ~~is~~ the water is heated the water near the flame ~~the~~ gets hot. Hot water raises up. The cold water from the sides move down towards the source of heat & this water also get hot and raises and the water from the sides moves down. This process continues till the whole water gets heated. eg :- The mode of heat transferred here is known as convection. eg :- heating of water in the container & transfer of heat by ~~conduction~~ ^{convection} in air.

~~Q16~~ (3) Radiation : The transfer of heat by radiation does not require any medium. It can take place whether the medium is present or not. eg :- from the sun the heat comes to us by the process of radiation.

Q16. Describe in brief about sea breeze & land breeze along with a diagram (fig 4.11)
 or

How does land and sea breeze occur?

A: The people living in coastal areas experience an interesting phenomena. (During the day the land gets hot faster than the sea. The air over the land becomes hotter & rises up. The cooler air from the sea rushes towards the land to take its place. The warm air from the land moves towards the sea to complete the cycle. The air from the sea is called as sea breeze.) (At night) it is exactly the reverse. (The water cools down more slowly than the land so the cool air from the land moves towards the sea this is called land breeze.)

Q17. Define Heat.

A: Heat is a transfer of kinetic energy from one medium or object to another or from an energy source to a medium or an object. Such energy transfer can occur in three ways that is conduction, convection & radiation.

Q18. Define Temperature.

A: Temperature is a measure of degree of hotness of an object.

Q19. Def what is a celsius scale?

A: It is an unit of the temperature of an object.

Q20. What are insulators?

A: The materials that do not allow heat to pass through them are called insulators.
eg:— rubber & wood.

Q21. Define Radiation?

A: It is a method of transfer of heat which requires no medium. eg:— sunlight reaches us through the process of radiation.

Q22 Why is a normal body temperature taken as 37°C even though it is not the same for every individual?

A: The normal body temperature is taken as 37°C because it is the average body temperature for a large no. of healthy persons. The temperature of the human body ^{does not go} below 35°C and above 42°C .

Q23 Write 2 examples of conductors of heat.

A: Iron rod & copper wire.

Q24 Write 2 examples of insulators of heat.

A: Mica, wood & plastic.

Q25 Why is it comfortable to wear light colour clothes in summer & dark colour clothes in winter?

A: Dark coloured ^{clothes} absorb radiation better than ~~make~~ light colour clothes. That is the reason we feel more comfortable in light

Chapter - 5 Acids, Bases and Salts

v. gmp

Name of acids	Found in
Acetic acid	Vinegar
Formic acid	Ant's sting
Citric acid	Citrus fruits such as oranges, lemon, etc.
Lactic acid	Curd
Oxalic acid	Spinach
Ascorbic acid (Vitamin C)	Amla, Citrus fruits
Tartaric acid	Tamarind, grapes, unripe mangoes, etc.

All the acids mentioned above occur in nature.

Name of acids	Found in
Calcium hydroxide	Limewater
Ammonium hydroxide	Window cleaner
Sodium hydroxide/ Potassium hydroxide	Soap
Magnesium hydroxide	Milk of magnesia

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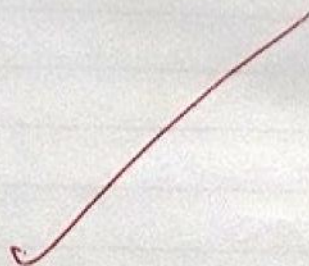
Red Litmus Paper



Blue Litmus Paper



Fig - 5.1 (b) Red and Blue litmus paper



Q1. Learn the following table thoroughly

Q1. Write the table given on pg-50 and learn thoroughly.

Q2. Define acids

or

acidic substances or acidic solution.

A: Acids are substances that turn blue litmus paper into red. The nature of these substances are called acidic and the solution is called as acidic solution.
example:- vinegar (acetic acid)

Q3. What are bases or basic substances or basic ^{basic} solution?

A: Bases are substances which turn red litmus into blue. The nature of such substances is called basic and the solution is called as basic solution. eg:-
baking soda (~~sodium bicarbonate~~)

Q4. What are indicators?

A: Indicators are special types of substances that are used to test whether a substance is acidic or basic ~~or~~ in nature. eg:- turmeric ~~and~~ litmus.

Q5. What do ^{you} mean by neutral solution?

A: The solution which do not change the colour of ^{litmus} ~~either~~ red or blue ~~litmus~~ are known as neutral solution. These substances are neither acidic nor basic.

Q6. Ammonia ^{hydroxide} is found in many household product, such as window cleaner. It turns red litmus blue.

A: ~~Ammonia~~ is found in many household

A: Basic.

Q7. Write any 3 differences between acids and bases.

Acids

Bases

①. Acids are sour in taste.

② Bases are bitter in taste.

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②. Acids are not soapy on touching

② Bases are soapy on touching.

③ Acids turn blue litmus paper into red.

③ Bases turn red litmus paper into blue.

Q8. Name the type of acid found in curd.

A: Lactic acid.

Q9. What is acid ^{raining} weathering? Where does this acid come from? Also state the damaging effects of acid rain.

A: The rain containing ^{excess} acids is called acid rain. The rain becomes acidic because of carbon-dioxide, sulphur dioxide and nitrogen dioxide dissolve in rain drops to form carbonic acid, sulphuric acid and nitric acid respectively. They come from factories, vehicles, etc as pollutants in the air. It can cause great damage to the buildings, historical monuments, plants & animals.

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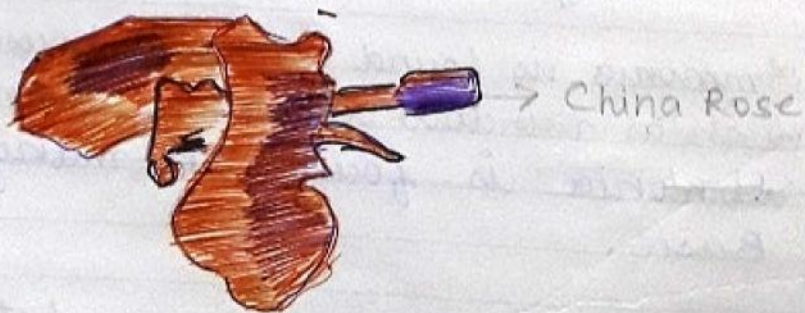
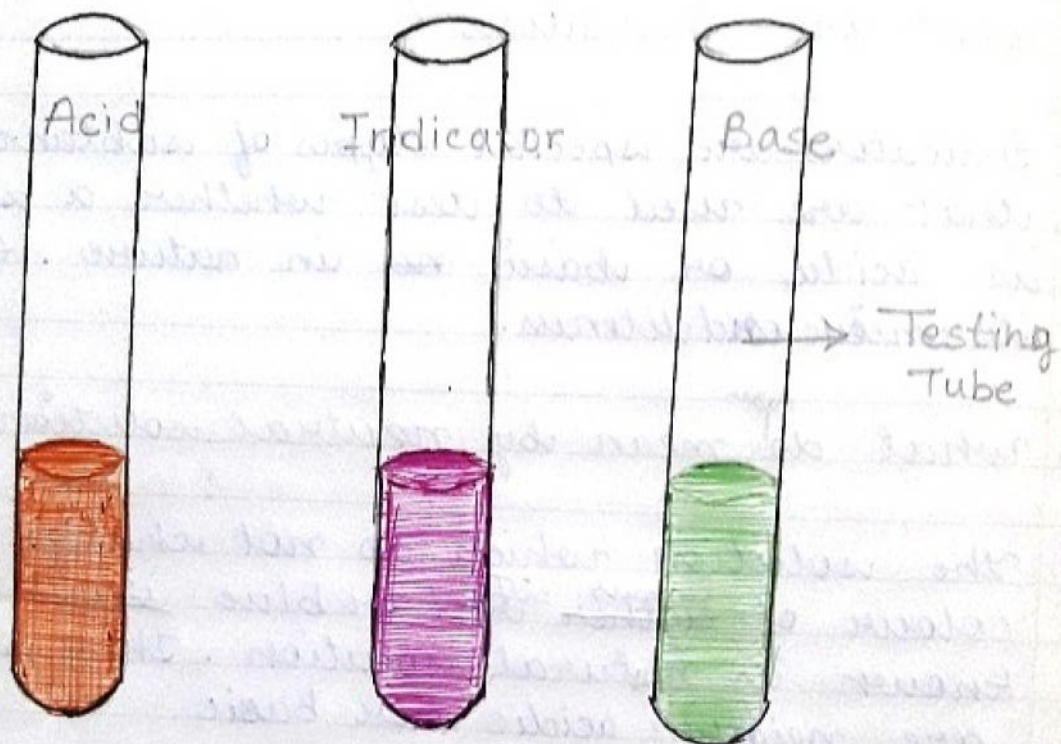


Fig. 5.3 China rose flower and indicator prepared from it

Q10. Is the distilled water acidic, basic or neutral? ^{neutral} How would you ~~not~~ verify it?

A: (1) Distilled water does not turn blue litmus red, therefore it is not acidic.

(2) Distilled water does not turn red litmus blue. Therefore, it is not basic.

(3) Therefore distilled water is neutral in nature.

Q11. Blue litmus paper is dipped in a solution. It remains blue. What is the nature of the solution?

A: Acids turn blue litmus red and bases turn red litmus blue because there is no colour change in the blue litmus paper it ~~is~~ can be a basic solution or neutral solution. Put a few drops of the ~~solution~~ on red litmus paper. If it turns blue, it is a basic in nature & if no colour change is there it is a neutral solution.

Q12. Deyi has a few bottles of soft drink in his restaurant but unfortunately these are not labeled. He has to serve the drinks on the demand of customer. One customer wants acidic drink another want

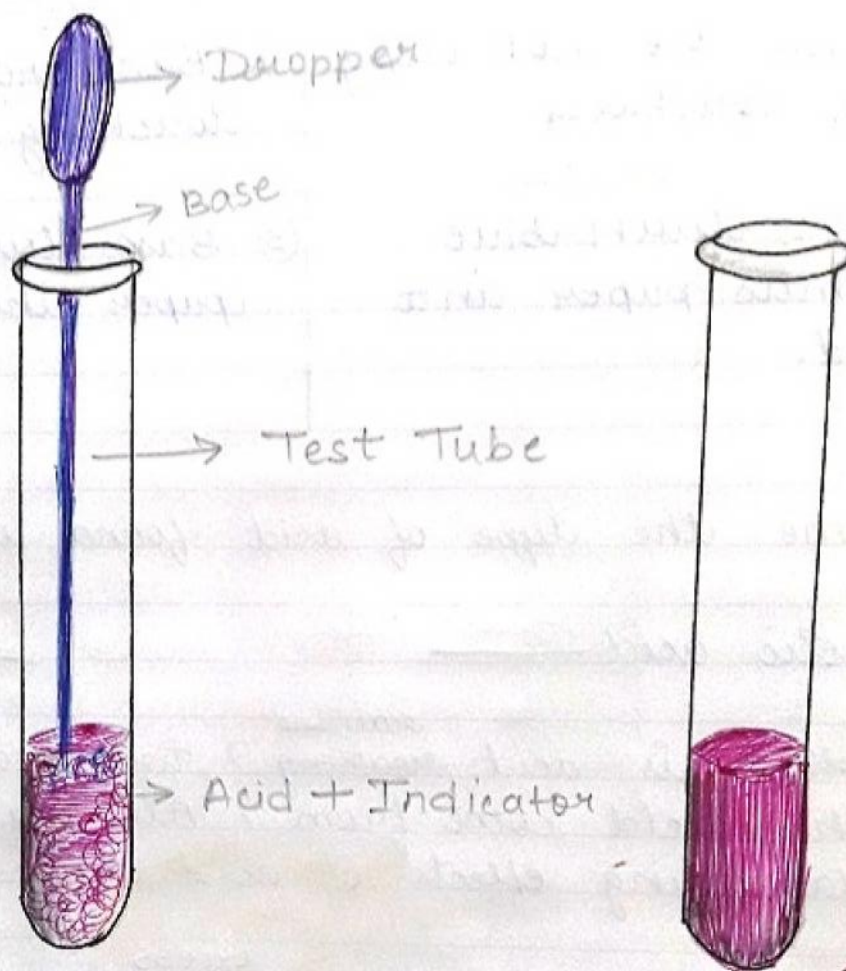


Fig 5.4. Process of neutralisation.

basic and the third one want natural drink. How will Doryi decide which drink is to be served to whom?

A: (1) When the drinking solution turns blue litmus red it is an acidic drink.

(2) When the drinking solution turns red litmus blue it is a basic drink.

(3) When the drinking solution does not change ^{the} colour either with ~~or~~ red or blue litmus paper it is a [^]neutral drink.

Q13 Name the source from which litmus solution is obtained. What is the ~~use~~ use of this solution?

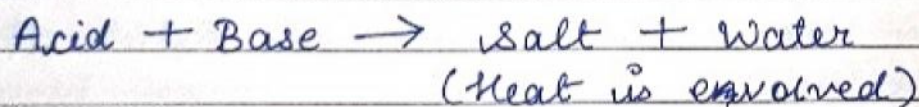
A: Litmus solution is obtained from lichens. The solution has a purple colour in distilled water. It turns into red if the solution is acidic and blue if it is basic. It is used to test whether a substance is acidic or basic. It is an indicator. It changes its colour when added to acidic or basic solution.

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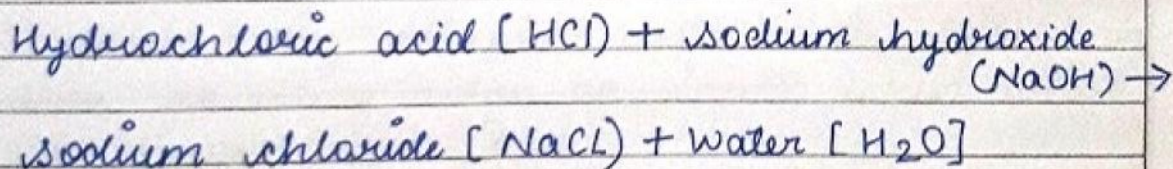
Q14. Describe the process of neutralisation with an example. What is produced during this process.

A: The reaction between an acid and a base is known as neutralisation. Salt & water are produced during this process with the evolution of heat.

(Equation)



* The following reaction is an example:



Q15. How can china rose be used as an indicator? Describe the method.

A: Collect some china rose (gudhal) petals & place them in a beaker. Add some warm water. Keep the mixture for some time till the water becomes coloured. Use the coloured water as an indicator on

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shampoo, lemon juice and vinegar. China rose indicates the acidic solution. If the colour is dark pink (magenta) and basic solution if it is green.

Q16. Explain why antacid tablet is taken when you suffer from indigestion?

A: Our stomach contains hydrochloric acid. It helps in digestion but too much of acid causes indigestion. So we take antacid tablets because it neutralises the effect of excessive acid.

Q17. Explain why calamine solution is applied on the skin to sting when an insect skin when an insect bites?

A: When an insect bites, it injects the acidic liquid (formic acid) that is formic acid into the skin. The effect of the acid can be neutralised by rubbing moist baking soda or calamine solution which contains zinc carbonate.

Q18 Explain why factory waste is neutralised before disposing in it into the water bodies?

A: The waste of many factories contain acids. If they are allowed to flow into the water bodies the acid will kill ^{the} fish and other organisms. The factory wastes are therefore neutralised by adding basic solution.

Q19. what will happen ~~not~~ if ~~turmeric~~ turmeric is added to acidic, basic and natural solution?

A: (1) When turmeric powder is added to acidic solution then its colour is yellow.

(2) When turmeric powder is added to basic solution then its colour is reddish-brown.

(3) When turmeric powder is added to natural solution then its ~~colour~~ is. There is no change in ^{the} colour.

Quick lime

Q20 (Slaked lime) is added to ~~an~~ acidic soil and organic matter is added to basic ^{soil} solution. why?

A: Plants do not grow well when the soil is too much acidic or basic in nature. To ~~return~~ neutralise the acid in the soil quick lime is added. Organic matter releases acid which neutralises the basic nature of the soil.

Q21 Three liquids are given to you. One is HCl and the other NaOH and another is sugar solution. How will you identify them. you have only turmeric indicator.

A: (1) When turmeric powder is added to acidic solution (HCl) the colour becomes yellow.

(2) When turmeric powder is added to ~~basic~~ basic solution (NaOH) the ~~a~~ colour becomes reddish-brown.

(3) When turmeric powder is added to sugar solution there is no change in ~~colour~~ colour.

Teacher's Signature _____

Date 18/7/18

Expt. No. _____

Page No. 68

Chapter - 6

Physical and Chemical changes.

Q1. * Define the following :-

1. Chemical reaction - The change in which one or more new ^{substances} ~~substances~~ are formed is called chemical change. A chemical change is also called a chemical reaction.
2. Chemical Property - A chemical property is a property or behaviour of a substance when it ~~under ground~~ undergoes chemical change or reaction. Chemical properties of a substance can include toxicity, oxidation and flammability. Iron ~~oxide~~ oxidises and forms rust.
3. Physical changes - A change in which a substance undergoes ~~for~~ a change in its physical property is called physical change.
4. Chemical change - A change in which one or more new substances are formed is called chemical change.

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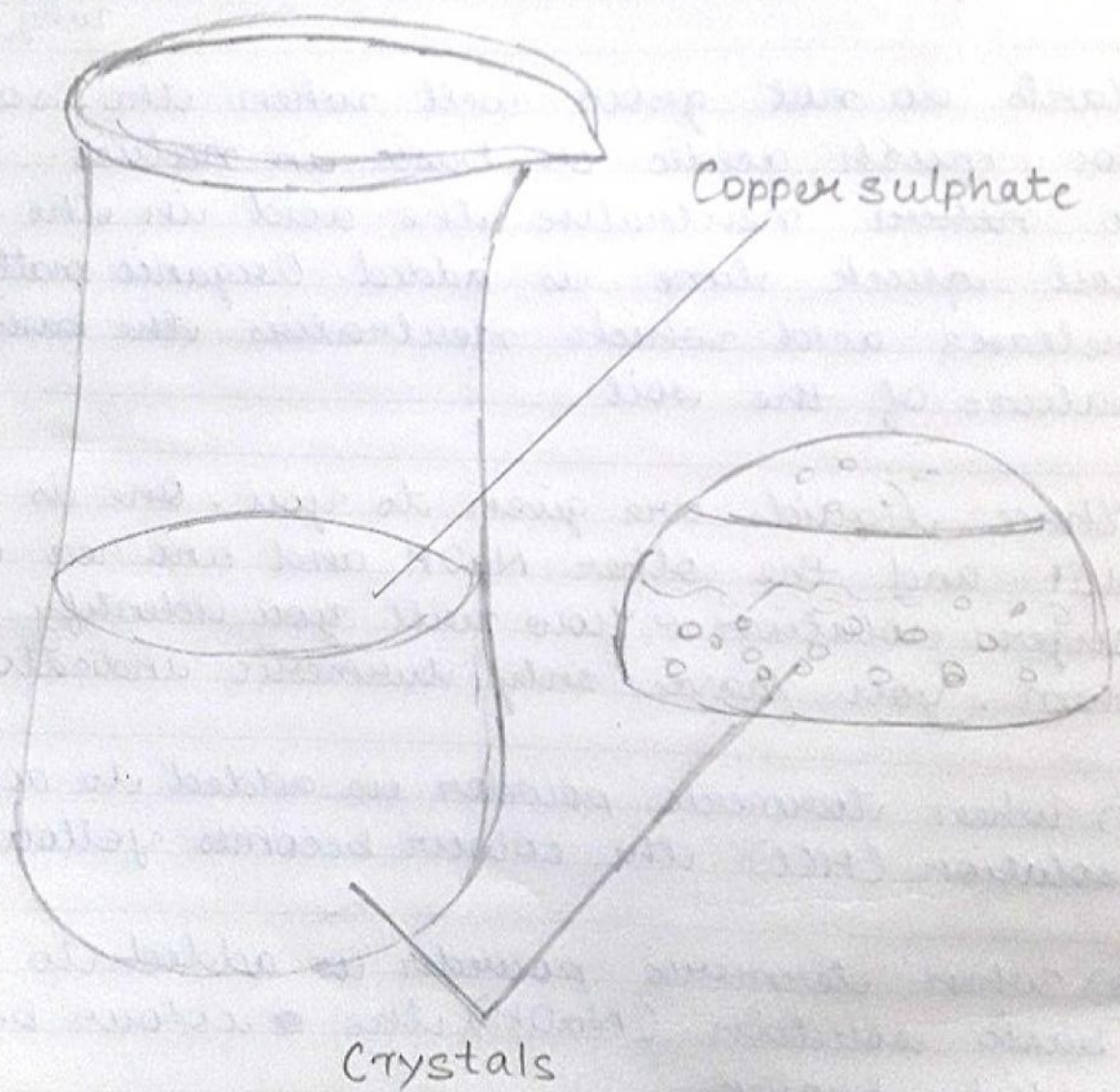


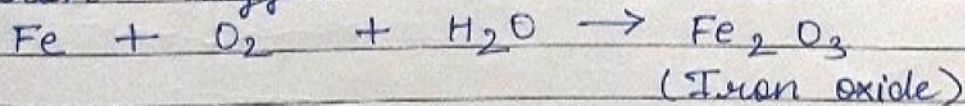
Fig 6.6 — Crystals of copper sulphate.

5. Physical property - Property such as shape, size, colour and state of a substance are called as physical properties.

Q2. Explain the process of rusting of Iron. Give the equation also.

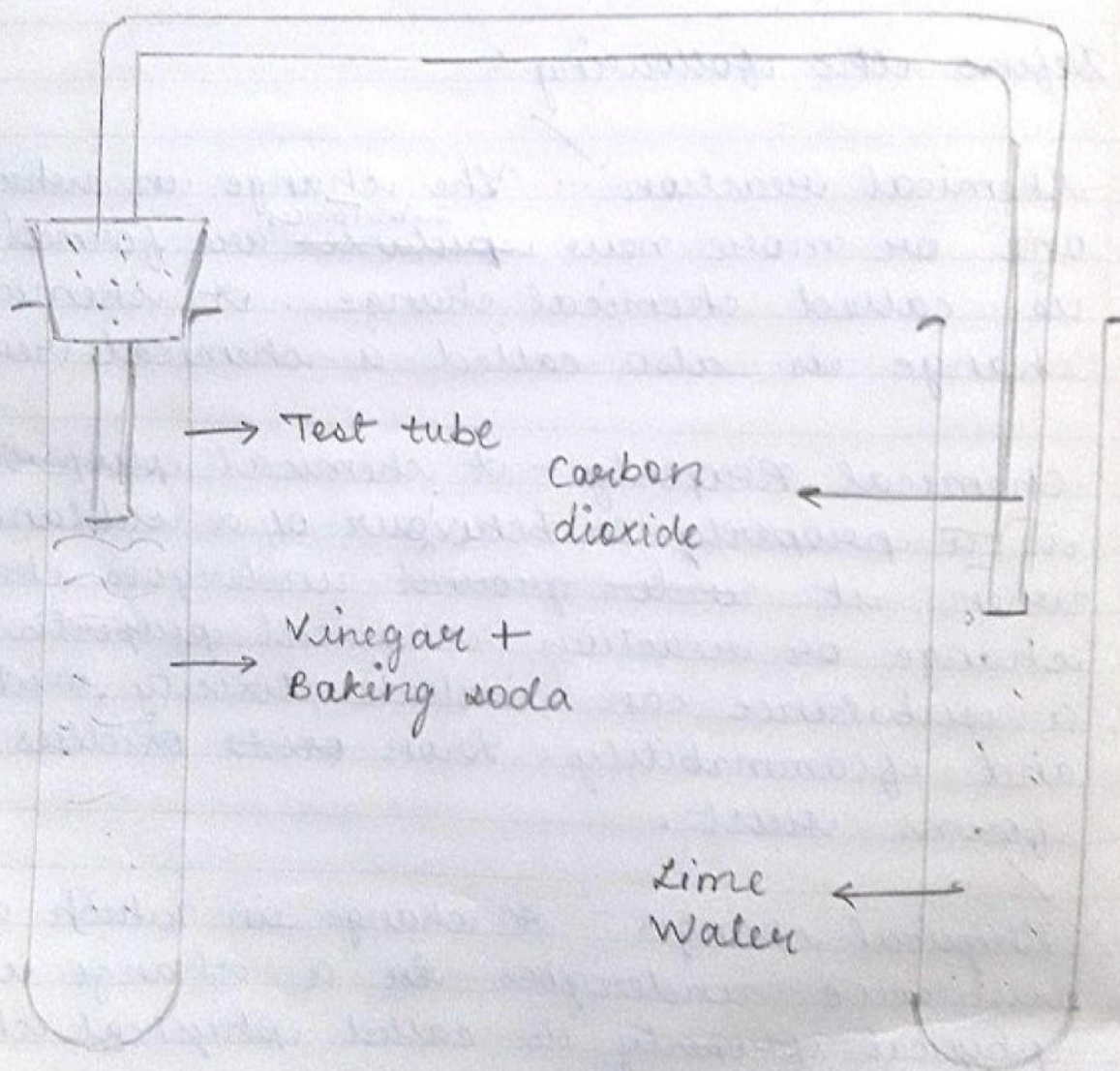
A: When we leave a piece of Iron in the open for sometime it ^{acquires} a film of brownish substance. This substance is called rust and this process is called rusting. For rusting the presence of water (water vapour), oxygen is essential. In fact if the content of moisture is high which means if it is more humid the rusting will be faster. The process of rusting can be represented by the following equation :-

Iron + Oxygen + water $\rightarrow$ Rust.



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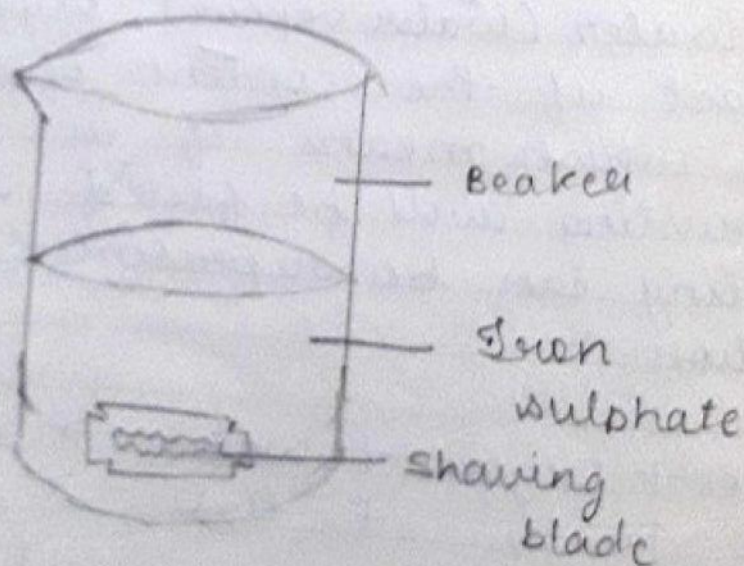
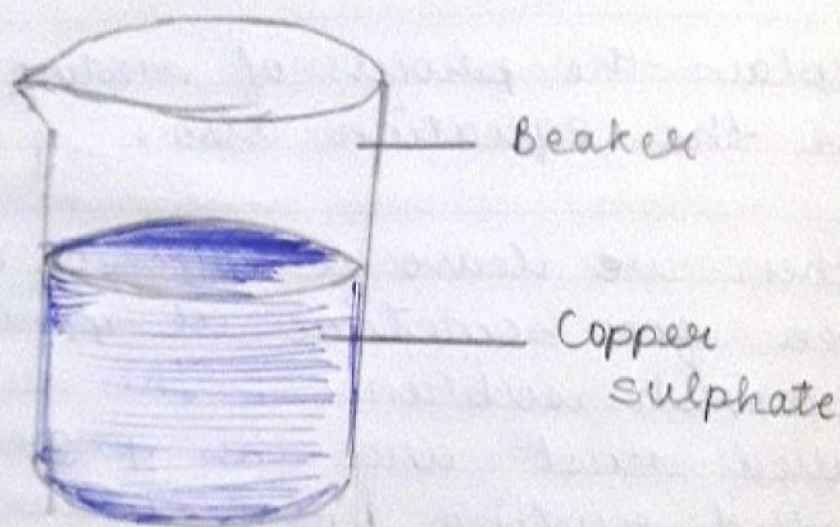
Set up to pass gas through lime water

Q3. Write a note on 'Various methods used to prevent rusting'.

A: We should prevent Iron articles from coming in contact with the oxygen or water or both. One simple way is to apply a coat of paint or grease. In fact these coats should be applied regularly to prevent rusting. Another way is to deposit a layer of metal like zinc or chromium or iron. This process of depositing a layer of ^{zinc} or ^{chromium} is called Galvanisation. The iron pipes we use in our homes to carry water are galvanised to prevent rusting.

Q4. What is meant by crystallisation?

A: Large crystals of pure substances can be formed from their solution. This process is called crystallisation.



Change in colour of the copper sulphate solution due to reaction with zinc.

Q5. What is ~~stain-steel~~ stainless steel made up of?

A: Stainless steel is made up of mixing iron with carbon and ~~met~~ metals like chromium, nickel and manganese. It does not rust.

Q6. Explain why rusting of iron objects are faster ^{in coastal} areas than in deserts?

A: For rusting the presence of both oxygen and water or water vapour is essential. Coastal areas have a very humid weather where there is high humidity. Therefore, rusting is also much faster in coastal areas. Deserts are very hot and dry hence, rusting does not occur ^{here} there.

Q7. Explain how painting of an iron gate prevents it from rusting?

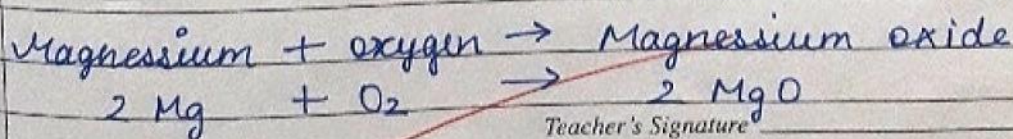
For rusting the presence of both oxygen and water (or water vapour) is essential. A coat of painting prevents iron articles from coming in contact with oxygen and water. In fact these coats of painting should be applied regularly to prevent rusting.

Q8. Explain ^{the} chemical ~~change~~ ^{what occurs} when a ribbon of magnesium is burnt ^{in air}. What happens when the product of the chemical reaction is mixed with water? Also write the relevant chemical equation.

A: Aim : To show the change when a ribbon of magnesium ^{is} burnt in air and dissolved in water.

Materials Required : ^{bb} a ribbon, a small piece of magnesium, sand paper, a match box, red and blue litmus paper.

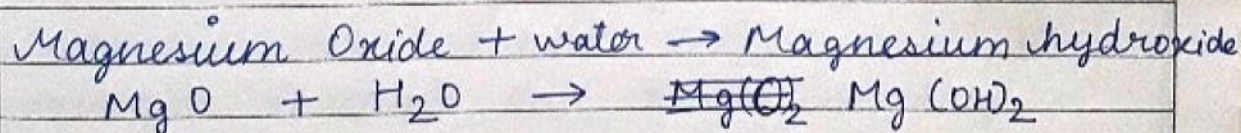
Method : (i) Get a small piece of the ribbon of magnesium and clean its tip with a sand paper. Bring the tip near a candle flame. It burns with a brilliant white light. When it is completely burnt it leaves behind a powdery ash. ~~The change can be represented by~~
The change can be represented by the following equation :-



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Collect the ash and mix it with a small amount of water. Stir the mixture well. Test the mixture with blue & red litmus solution.

Observation :- (1) The mixture turns red litmus into blue. There is no colour change in blue litmus solution. This shows that they are basic in nature. On dissolving the ash with water. It forms a new substance it changes red litmus into blue. This can be represented by the following equation :-



Conclusion : This activity shows how magnesium burns in air and produced magnesium oxide (MgO) which when dissolved in water produces magnesium hydroxide which is a basic substance and changes red litmus blue.

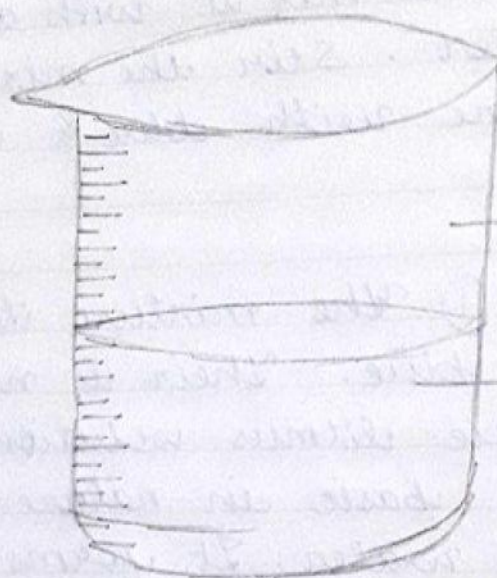
Q9. Explain why the burning of wood and cutting it down into small pieces of wood are considered as two different types of changes.

A: Burning of wood is a chemical change from which a new substance ^{ash} is formed which can not be reversed (irreversible change). Cutting of wood is a physical change and as there is a change only in the size and shape of the ~~write~~ object wood.

Q10. With the help of an activity explain the chemical reaction taking place when ~~the~~ ^{an iron} nail or a used shaving blade is put into copper sulphate solution. Include the chemical equation involved and draw the diagram.

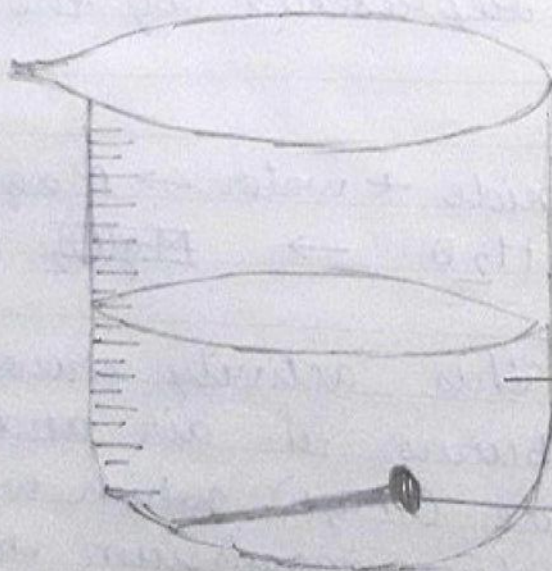
A: Aim: To show the chemical change that takes place when a nail or a used shaving blade is put into copper sulphate solution (CuSO_4).

Materials Required: Nail or a used shaving blade, copper sulphate ^esolution, dilute sulphuric acid, 2 beakers, test tube.



→ Beaker.

→ Copper sulphate



→ Iron sulphate

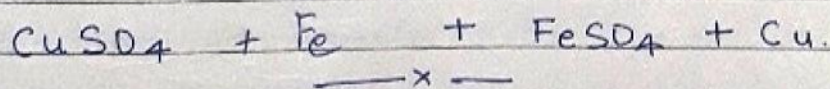
→ Iron nail

Change in the colour of copper sulphate solution due to reaction with iron.

Method :- Dissolve about a teaspoon full of copper sulphate solution (blue vitriol) in about half of a cup of water in a glass tumbler or a beaker. Add a few drops of dilute sulphuric acid to the solution. We will get blue coloured solution. Save a small sample of solution in a test tube or a small glass bottle. Drop a nail or a used shaving blade into the remaining solution. Wait for half an hour. And observe the colour of the solution.

Observation : The change of the colour of the solution its blue to green is due to the formation of iron sulphate as new substance. The brown deposit on the ~~copper~~^{iron} nail is copper and other new substances. Equation :-

Copper sulphate solution + Iron $\rightarrow$ Iron sulphate + Copper.



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conclusion: This activity shows that a chemical reaction takes place when an ^{iron} nail or a ~~used~~ ~~shaving~~ blade is put into a copper sulphate solution.

Q11. Explain ~~crystallization~~ crystallization section with an activity and also give an example.

OR

Describe how crystals of copper sulphate are prepared.

A: Aim: To show how ~~crystallization~~ crystals of copper sulphate (CuSO_4) are formed.

Materials required: Water, beaker, sulphuric acid (H_2SO_4) copper sulphate (CuSO_4) powder.

Method: Take a cup full of water and in a beaker and add a few drops of dilute sulphuric acid. ~~Keep~~ the water heat the water. When it starts boiling add copper sulphate powder till no more ^{powder} can be dissolved. Filter the solution and allow it to cool. Do not disturb when it is cooling. Look at the solution after sometime.

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* Observation : We will find copper sulphate powder has changed into crystals.

* Conclusion : This activity shows that the crystals of copper sulphate are formed by the method of crystallisation.

Q12 Explain the chemical change which occurs when baking soda is mixed with vinegar. Write the ~~ex~~ equation and also draw the diagram.

Aim : To show a chemical change when baking soda is mixed with vinegar.

Materials Required : Two test tubes, powdered baking soda, V shaped ~~test~~ tube & freshly prepared lime water.

Method : Take about a tea spoon full of vinegar in a test tube. Add a pinch of baking soda into it. We would hear a hissing sound and see bubbles of gas coming out. Pass this through the freshly prepared lime water.

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* Observation:- The change in the test tube is as follows:-

Vinegar + Baking soda $\rightarrow$ Carbon-dioxide + other substances.

(Acidic acid) + sodium hydrogen carbonate $\rightarrow$ Calcium Carbon dioxide (CO_2) + other substances.

The reaction between carbon dioxide and lime water is as follows:-

Carbon dioxide (CO_2) + Limewater $\rightarrow$ Calcium carbonate + ^{water}
 $\text{CO}_2 + \text{Ca(OH)}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$

When carbon dioxide is passed through limewater, calcium carbonate is formed which makes lime water milky.

* Conclusion: This activity shows that a chemical ^{change} occurs when baking soda ^{mixes} with vinegar.

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Q.13. Explain how painting of an iron gate prevents it from rusting.

A: Rusting can be prevented by keeping the metals away from air and moisture by painting the iron gate we prevent the contact of atmospheric air & moisture with the metal iron so it prevents rusting.

Q.14. When baking soda is mixed with lemon juice bubbles are formed with the evolution of gas. What type of change is it? Explain.

A: When baking soda is mixed with lemon juice ^{bubbles} are formed and carbon dioxide gas is evolved. It is a chemical change.

Q.15. Explain how painting of an iron gate prevents it from rusting classify the following changes as physical or chemical.

- A:
- (i) ~~Photo~~ Photosynthesis - Chemical change
 - (ii) Dissolving sugar in water - Physical change.
 - (iii) Burning of coal - Chemical change

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(iv) Melting of wax - Physical change.

(v) Beating aluminium to make aluminium foil - Physical change.

(vi) Digestion of food - chemical change.

Q16. When a candle burns both physical & chemical change takes place why? Give another eg. of the if similar process in which both the physical & chemical changes takes place.

A: When a candle burns it melts to liquid and gases form which is a physical change. While burning it produces a new compound called ash from which a candle can not be obtained therefore, it is a chemical change and another where both physical and chemical change is taking place is making of bread from maida by using yeast.

Q17. How do you show that setting of curd is a chemical change?

A: When curd is formed from fresh milk the lactose sugar present in the milk

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is converted into lactic acid. Lactic acid ^{bacillus} helps in the conversion of milk into curd. The curd cannot be reversed back into the previous form that is milk therefore, setting of curd is a chemical change.

Q18. What is ~~gal~~ galvanisation?

A: The process of depositing a layer of zinc on iron is called galvanisation. eg:- The iron pipes used in our home for carrying water are galvanised to prevent rusting.

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(continued in the next copy)